TOPPERZ @ WORK EDUCATION CENTRE

CLASS: CBSE XI TOPIC: s - BLOCK ELEMENTS

Time: 45 MINUTES

1. Name the elements present in Group 2 [2]

2. The atomic radii of alkaline earth metals are smaller than those of the corresponding alkali metals. Explain why? [2]

3. Why do alkaline earth metals have low ionization enthalpy? [1]

4. The second ionization enthalpy of calcium is more than the first. How is that calcium forms $CaCl_2$ and not CaCl give reasons. [2]

5. State one reason for alkaline earth metals in general having a greater tendency to form complexes than alkali metals. [1]

6. Name the metal amongst alkaline earth metals whose salt do not impart colour to a non-luminous flame. [2]

7. Compounds of alkaline earth metals are more extensively hydrated than those of alkali metals. Give reason. [1]

8. The melting and boiling points of alkaline metals are higher than alkali metals. Give reason. [1]

9. Which member of the alkaline earth metals family has: [2]

   (i) least reactivity (ii) lowest density (iii) highest boiling point

   (iv) maximum reduction potential

10. The alkaline earth metals are called s – block elements. Give reasons. [2]

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1. What is the nature of oxide formed by Be? [1]

2. Why does beryllium show similarities with Al? [1]

3. Why is Calcium preferred over sodium to remove last traces of moisture from alcohol? [2]

4. Why is beryllium carbonate unusually unstable thermally as compared to the other carbonates of this group? [1]

5. Name the metal amongst alkaline earth metals whose salt do not impart colour to a non – luminous flame. [2]

6. Why sulphates of Mg and Be soluble in water? [1]

7. Why does the solubility of alkaline earth metal hydroxides in water increase down the group? [2]

8. Why beryllium is not attacked by an acid easily? [1]

9. Give the reaction of magnesium with air? [2]

10. Beryllium is reducing in nature. Why? [2]